

ATUL CLASSES

Test / Exam Name: Atul Classes	Standard: 10th	Subject: Science
Student Name: _____	Section: _____	Roll No.: _____
Questions: 100		Time: 01:00 hh:mm
Negative Marks: 0		Marks: 100

- Q1. Consider the following oils:

1. Mobil oil.
2. Castor oil.
3. Turpentine oil.
4. Kerosene.
5. Mustard oil.
6. Coconut oil.

Which of these can be used for preparation of soap?

A I, II, III, VI B II, V, VI C II, III, V, VI D II, III, VI

1 Mark
- Q2. $\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$ is:

A Washing soda B Baking soda C Bleaching powder D Tartaric acid

1 Mark
- Q3. Fresh milk has a pH of 6. When milk changes into curd, the pH value will:

A Become 7 B Become less than 6 C Become more than 7 D Remain unchanged.

1 Mark
- Q4. Aqueous solution of copper sulphate reacts with aqueous ammonium hydroxide solution to give.

A Brown precipitate. B Pale blue precipitate. C White precipitate. D Green precipitate.

1 Mark
- Q5. In which of the following bleaching powder is not used?

A For bleaching wood pulp. B For water sterilisation. C As an oxidising agent.
D For sugar decolourisation.

1 Mark
- Q6. The colour of methyl orange indicator in a solution is yellow. The pH of this solution is likely to be:

A 7 B Less than 7 C 0 D More than 7

1 Mark
- Q7. Plaster of Paris hardens by:

A Giving off CO_2 B Changing into C Combining with water CaCO_3
D Giving out water

1 Mark
- Q8. Calcium phosphate is present in tooth enamel, its nature is:

A Basic B Acidic C Neutral D None of the above

1 Mark
- Q9. You are having five solutions A, B, C, D and E with pH values as follows:
A = 1.8, B = 7, C = 8.5, D = 8 and E = 5
Which solution would be most likely to liberate hydrogen with magnesium powder?

A Solution A and B. B Solution A. C Solution C. D All of the above.

1 Mark
- Q10. If you take some distilled water in a test-tube, add an equal amount of acetic acid to it, shake the test-tube well and leave it undisturbed on the test-tube stand, then after about 5 minutes, what would you observe?.

A There is a layer of water over the layer of acetic acid.
B A precipitate is settling at the bottom of the test-tube.
C Bubbles of colourless gas are coming out of the test-tube.
D There is a clear, colourless transparent solution in the test-tube.

1 Mark
- Q11. Which of the following is a weak acid?

A Acetic acid B Hydrochloric acid C Sulphuric acid D Nitric acid

1 Mark

Q12. If a few drops of a concentrated acid accidentally spills over the hand of a student, what should be done? **1 Mark**

- A** Wash the hand with saline solution.
- B** Wash the hand immediately with plenty of water and apply a paste of sodium hydrogen carbonate.
- C** After washing with plenty of water apply solution of sodium hydroxide on the hand.
- D** Neutralise the acid with a strong alkali.

Q13. A solution turns red litmus blue. Its pH is likely to be: **1 Mark**

- A** 1
- B** 4
- C** 5
- D** 10

Q14. When copper oxide and dilute hydrochloric acid react, colour changes to: **1 Mark**

- A** White
- B** Bluish-green
- C** Blue-black
- D** Black

Q15. Bee sting contains: **1 Mark**

- A** An acidic liquid.
- B** A salt solution.
- C** An alkaline liquid.
- D** An alcohol.

Q16. For question two statements are given-one labelled Assertion (A) and the other labelled Reason (R). **1 Mark**

Select the correct answer to these questions from the codes (a), (b), (c) and (d) as given below:

Assertion: pH of ammonium nitrate solution is acidic.

Reason: Solution of a salt of weak base and strong acid is acidic.

- A** Both A and R are true, and R is correct explanation of the assertion.
- B** Both A and R are true, but R is not the correct explanation of the assertion.
- C** A is true, but R is false.
- D** A is false, but R is true.

Q17. You are asked to prepare hard water in your laboratory. Select a group from the following groups of salts, any salt of which you may dissolve in distilled water to obtain hard water. **1 Mark**

- A** NaCl; Na₂SO₄; KCl
- B** NaCl; CaCl₂; KCl
- C** CaCl₂; CaSO₄; MgSO₄
- D** Na₂SO₄; CaSO₄; MgSO₄

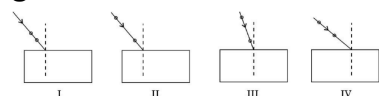
Q18. Complete the reaction: $2\text{NaOH} + \text{Zn} \rightarrow \text{_____} + \text{H}_2$. **1 Mark**

- A** Na ZnO₂
- B** Na₂ ZnO₃
- C** Na₂ ZnO₂
- D** ZnO

Q19. The chemical formula for plaster of Paris is: **1 Mark**

- A** CaSO₄ · 2H₂O.
- B** CaSO₄ · H₂O.
- C** CaSO₄ · $\frac{1}{2}$ H₂O.
- D** 2CaSO₄ · H₂O.

Q20. Select from the following the best experimental set-up for tracing the path of a ray of light through a glass slab: **1 Mark**



- A** I
- B** II
- C** III
- D** IV

Q21. For question two statements are given-one labelled Assertion (A) and the other labelled Reason (R). **1 Mark**

Select the correct answer to these questions from the codes (a), (b), (c) and (d) as given below:

Assertion: Common salt is used for the preparation of many chemicals such as sodium hydroxide, bleaching powder, baking soda, washing soda etc.

Reason: Main source of sodium chloride is sea water.

- A** Both A and R are true, and R is correct explanation of the assertion.
- B** Both A and R are true, but R is not the correct explanation of the assertion.
- C** A is true, but R is false.
- D** A is false, but R is true.

Q22. Identify the correct representation of reaction occurring during chloralkali process. **1 Mark**

- A** $2\text{NaCl(l)} + 2\text{H}_2\text{O(l)} \rightarrow 2\text{NaOH(l)} + \text{Cl}_2\text{(g)} + \text{H}_2\text{(g)}$
- B** $2\text{NaCl(aq)} + 2\text{H}_2\text{O(aq)} \rightarrow 2\text{NaOH(aq)} + \text{Cl}_2\text{(g)} + \text{H}_2\text{(g)}$
- C** $2\text{NaCl(aq)} + 2\text{H}_2\text{O(l)} \rightarrow 2\text{NaOH(aq)} + \text{Cl}_2\text{(aq)} + \text{H}_2\text{(aq)}$
- D** $2\text{NaCl(aq)} + 2\text{H}_2\text{O(l)} \rightarrow 2\text{NaOH(aq)} + \text{Cl}_2\text{(g)} + \text{H}_2\text{(g)}$

Q23. A student takes 2 mL acetic acid in a dry test tube and adds a pinch of sodium hydrogen carbonate to it. He makes the following observations: **1 Mark**

1. A colorless and odourless gas evolves with a brisk effervescence.
2. The gas turns lime water milky when passed through it.
3. The gas burns with an explosion when a burning splinter is brought near it.
4. The gas extinguishes the burning splinter that is brought near it.

The correct observations are:

- A** I, II and III **B** II, III and IV **C** III, IV and I **D** IV, I and II

Q24. Egg shell is made up of: **1 Mark**

- A** CaCO_3 **B** CaO **C** Ca(OH)_2 **D** CaCl_2

Q25. Which of the following types of medicine is used for treating indigestion caused by over-eating? **1 Mark**

- A** Antibiotic. **B** Analgesic. **C** Antacid. **D** Antiseptic.

Q26. Which of the following is acidic in nature? **1 Mark**

- A** Lime juice. **B** Human blood. **C** Lime water. **D** Antacid.

Q27. A student takes about 6 mL distilled water in four test tubes marked P, Q, R and S. He dissolves sodium sulphate in P, potassium sulphate in Q, calcium sulphate in R and magnesium sulphate in S. After that he adds equal amount of soap solution in each test tube. On shaking these test tubes, he would observe a good amount of lather in the test tubes marked. **1 Mark**

- A** P and Q. **B** Q and R. **C** R and S. **D** P and S.

Q28. Match the chemical substances given in Column (A) with their appropriate application given in Column (B): **1 Mark**

	Column (A)		Column (B)
A.	Bleaching powder	1.	Preparation of glass
B.	Baking soda	2.	Production of H_2 and Cl_2
C.	Washing soda	3.	Decolourisation
D.	Sodium chloride	4.	Antacid

- A** A—2, B—1, C—4, D—3 **B** A—3, B—2, C—4, D—1 **C** A—3, B—4, C—1, D—2 **D** A—2, B—4, C—1, D—3

Q29. Plaster of Paris is prepared by heating one of the following to a temperature of 100°C . This is: **1 Mark**

- A** $\text{CaSO}_3 \cdot 2\text{H}_2\text{O}$ **B** $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$ **C** $\text{CaCO}_3 \cdot 2\text{H}_2\text{O}$ **D** $\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$

Q30. For question two statements are given-one labelled Assertion (A) and the other labelled Reason (R). Select the correct answer to these questions from the codes (a), (b), (c) and (d) as given below: **1 Mark**

Assertion: Baking soda is prepared by chlor-alkali process.

Reason: Brine decomposes to sodium hydroxide on passing electricity through it.

- A** Both A and R are true, and R is correct explanation of the assertion.
B Both A and R are true, but R is not the correct explanation of the assertion. **C** A is true, but R is false.
D A is false, but R is true.

Q31. Raisins are wiped off gently before final weighing with help of: **1 Mark**

- A** A filter paper. **B** A cotton piece. **C** A cloth piece. **D** A polyethene piece.

Q32. Tomato is a natural source of which acid? **1 Mark**

- A** Acetic acid **B** Citric acid **C** Tartaric acid **D** Oxalic acid

Q33. One of the constituents of baking powder is sodium hydrogencarbonate, the other constituent is: **1 Mark**

- A** Hydrochloric acid. **B** Tartaric acid. **C** Acetic acid. **D** Sulphuric acid.

Q34. Nettle sting is a natural source of which acid? **1 Mark**

- A** MetiWanoic acid **B** Lactic acid **C** Citric acid **D** Tartaric acid

Q35. What happens when a solution of an acid is mixed with a solution of a base in a test tube? **1 Mark**

1. The temperature of the solution increases.
2. The temperature of the solution decreases.
3. The temperature of the solution remains the same.
4. Salt formation takes place.

A 1 only **B** 1 and 3 **C** 2 and 3 **D** 1 and 4

Q36. Which of the following is not a mineral acid? **1 Mark**

A Hydrochloric acid. **B** Citric acid. **C** Sulphuric acid. **D** Nitric acid.

Q37. Which of these give out hydronium ions in water? **1 Mark**

A Acids **B** Base **C** Both A and B **D** None of these

Q38. If a few drops of a concentrated acid accidentally spills over the hand of a student, what should be done? **1 Mark**

- A** Wash the hand with saline solution.
B Wash the hand immediately with plenty of water and apply a paste of sodium hydrogencarbonate.
C After washing with plenty of water apply solution of sodium hydroxide on the hand.
D Neutralise the acid with a strong alkali.

Q39. Consider the following oils: **1 Mark**

1. Mobil oil.
2. Castor oil.
3. Turpentine oil.
4. Kerosene.
5. Mustard oil.
6. Coconut oil.

Which of these can be used for preparation of soap?

A I, II, III, VI **B** II, V, VI **C** II, III, V, VI **D** II, III, VI

Q40. Which of the following can be used in the laboratory to test whether a solution is a base or an acid? **1 Mark**

A Olfactory indicator **B** Acid indicator **C** Basic indicator **D** Regular indicator

Q41. Toothpastes are ____ in nature. **1 Mark**

A Acidic **B** Basic **C** Neutral **D** None of these

Q42. If pH of solution is 13, it means that it is: **1 Mark**

A Weakly acidic **B** Weakly basic **C** Strongly acidic **D** Strongly Basic

Q43. A solution turns blue litmus red. Its pH is likely to be: **1 Mark**

A 7 **B** 5 **C** 8 **D** 14

Q44. The formula of baking soda is: **1 Mark**

A K_2CO_3 **B** $KHCO_3$ **C** $NaHCO_3$ **D** Na_2CO_3

Q45. For question two statements are given-one labelled Assertion (A) and the other labelled Reason (R). **1 Mark**

Select the correct answer to these questions from the codes (a), (b), (c) and (d) as given below:

Assertion: The chemical name of bleaching powder is calcium oxychloride.

Reason: Bleaching powder is used as an oxidising agent in chemical industries.

- A** Both A and R are true, and R is correct explanation of the assertion.
B Both A and R are true, but R is not the correct explanation of the assertion. **C** A is true, but R is false.
D A is false, but R is true.

Q46. Common salt besides being used in kitchen can also be used as the raw material for making: **1 Mark**

1. Washing soda.
2. Bleaching powder.
3. Baking soda.
4. Slaked lime.

A 1 and 2**B** 1, 2 and 4**C** 1 and 3**D** 1, 3 and 4

Q47. During the course of an experiment, 'to determine the percentage of water absorbed by raisins', raisins are weighed. **1 Mark**

- A** Every half an hour. **B** Every hour. **C** Once - only after completing the experiment.
D Two times - before soaking and after soaking for three hours.

Q48. Common salt besides being used in kitchen can also be used as the raw material for making: **1 Mark**

1. Washing soda.
2. Bleaching powder.
3. Baking soda.
4. Slaked lime.

- A** (i) and (ii) **B** (i), (ii) and (iv) **C** (i), (ii) and (iii) **D** (i), (iii) and (iv)

Q49. Which of the following is (are) true when HCl (g) is passed through water? **1 Mark**

1. It does not ionise in the solution as it is a covalent compound.
2. It ionises in the solution.
3. It gives both hydrogen and hydroxyl ion in the solution.
4. It forms hydronium ion in the solution due to the combination of hydrogen ion with water molecule.

- A** 1 only **B** 3 only **C** 2 and 4 **D** 3 and 4

Q50. For question two statements are given-one labelled Assertion (A) and the other labelled Reason (R). **1 Mark**

Select the correct answer to these questions from the codes (a), (b), (c) and (d) as given below:

Assertion: Antacids neutralize the effect of extra acid produced in the stomach during indigestion and thus provide relief.

Reason: Antacids are mild bases.

- A** Both A and R are true, and R is correct explanation of the assertion.
B Both A and R are true, but R is not the correct explanation of the assertion. **C** A is true, but R is false.
D A is false, but R is true.

Q51. Which of the following is treated with chlorine to obtain bleaching powder? **1 Mark**

- A** CaSO_4 **B** Ca(OH)_2 **C** Mg(OH)_2 **D** KOH

Q52. In terms of acidic strength, which one of the following is in the correct increasing order? **1 Mark**

- A** Water < Acetic acid < Hydrochloric acid. **B** Water < Hydrochloric acid < Acetic acid.
C Acetic acid < Water < Hydrochloric acid. **D** Hydrochloric acid < Water < Acetic acid.

Q53. The pH of the gastric juices released during digestion is: **1 Mark**

- A** Less than 7. **B** More than 7. **C** Equal to 7. **D** Equal to 0.

Q54. Baking soda is a mixture of: **1 Mark**

- A** Sodium carbonate and acetic acid. **B** Sodium carbonate and tartaric acid.
C Sodium hydrogen carbonate and tartaric acid. **D** Sodium hydrogen carbonate and acetic acid.

Q55. A salt whose aqueous solution will have a pH of more than 7 will be: **1 Mark**

- A** K_2CO_3 **B** K_2SO_4 **C** NaCl **D** NH_4Cl

Q56. Equal volumes of hydrochloric acid and sodium hydroxide solutions of same concentration are mixed and the pH of the resulting solution is checked with a pH paper. What would be the colour obtained? **1 Mark**

- A** Red. **B** Yellow. **C** Yellowish green. **D** Blue.

Q57. Which of the following salts will give an aqueous solution having pH of almost 7? **1 Mark**

- A** NH_4NO_3 **B** NH_4Cl **C** CaCl_2 **D** KCl

Q58. Sodium hydrogencarbonate when added to acetic acid evolves a gas. Which of the following statements are true about the gas evolved? **1 Mark**

1. It turns lime water milky.
2. It extinguishes a burning splinter.
3. It dissolves in a solution of sodium hydroxide.
4. It has a pungent odour.

A 1 and 2 **B** 1, 2 and 3 **C** 2, 3 and 4 **D** 1 and 4

Q59. Which of the has atmosphere made up of thick white and yellowish clouds of sulphuric acid? **1 Mark**

A Saturn **B** Mars **C** Venus **D** Uranus

Q60. Acids are _____ in taste. **1 Mark**

A Sour **B** Bitter **C** Spicy **D** Sweet

Q61. During the preparation of hydrogen chloride gas on a humid day, the gas is usually passed through the guard tube containing calcium chloride. The role of calcium chloride taken in the guard tube is to: **1 Mark**

A Absorb the evolved gas. **B** Moisten the gas. **C** Absorb moisture from the gas.
D Absorb Cl^- ions from the evolved gas.

Q62. Which one of the following salts does not contain water of crystallisation? **1 Mark**

A Blue vitriol **B** Baking soda **C** Washing soda **D** Gypsum

Q63. Which of the following are present in a dilute aqueous solution of hydrochloric acid? **1 Mark**

A $\text{H}_3\text{O}^+ + \text{Cl}^-$ **B** $\text{H}_3\text{O}^+ + \text{OH}^-$ **C** $\text{Cl}^- + \text{OH}^-$ **D** Unionised HCl.

Q64. Which one of the following is non-crystalline or amorphous? **1 Mark**

A Diamond **B** Graphite **C** Glass **D** Common Salt

Q65. The property which is not shown by acids is: **1 Mark**

A They have sour taste. **B** They feel soapy. **C** They turn litmus red.
D Their pH is less than seven.

Q66. A solution reacts with marble chips to produce a gas which turns lime water milky. The solution contains: **1 Mark**

A Na_2SO_4 **B** CaSO_4 **C** H_2SO_4 **D** K_2SO_4

Q67. Which one of the following types of medicines is used for treating indigestion? **1 Mark**

A Antibiotics **B** Analgesic **C** Antacid **D** Antiseptic

Q68. The chemical mostly used in the preparation of most of the soaps we use is. **1 Mark**

A Sodium chloride. **B** Potassium hydroxide. **C** Sodium hydroxide. **D** Potassium chloride.

Q69. Sodium hydroxide is used: **1 Mark**

A As an antacid **B** In manufacture of soap **C** As a cleansing agent **D** In alkaline batteries

Q70. Acetic acid solution turns: **1 Mark**

A Blue litmus red. **B** Red litmus blue. **C** Blue litmus colourless. **D** Red litmus colourless.

Q71. One of the following salts will give an alkaline solution on dissolving in water. This is: **1 Mark**

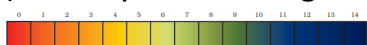
A Na_2CO_3 **B** Na_2SO_4 **C** NaCl **D** $(\text{NH}_4)_2\text{SO}_4$

Q72. A solution turns phenolphthalein indicator pink. The most likely pH of this solution will be: **1 Mark**

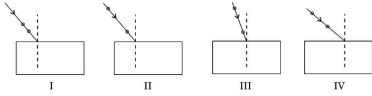
A 6 **B** 4 **C** 9 **D** 7

Q73. Equal volumes of hydrochloric acid and sodium hydroxide solutions of same concentration are mixed and the pH of the resulting solution is checked with a pH paper. What would be the colour obtained? **1 Mark**

(You may use colour guide given in Figure:



A Red. **B** Yellow. **C** Yellowish green. **D** Blue.

- Q74.** Sodium carbonate is a basic salt because it is a salt of: **1 Mark**
- A** Strong acid and strong base. **B** Weak acid and weak base. **C** Strong acid and weak base.
D Weak acid and strong base.
- Q75.** Which of the following statements is true for acids? **1 Mark**
- A** Bitter and change red litmus to blue. **B** Sour and change red litmus to blue.
C Sour and change blue litmus to red. **D** Bitter and change blue litmus to red.
- Q76.** For question two statements are given-one labelled Assertion (A) and the other labelled Reason (R).
 Select the correct answer to these questions from the codes (a), (b), (c) and (d) as given below:
Assertion: Strength of the acid or base decreases with dilution.
Reason: Ionization of an acid or a base increases with dilution. **1 Mark**
- A** Both A and R are true, and R is correct explanation of the assertion.
B Both A and R are true, but R is not the correct explanation of the assertion. **C** A is true, but R is false.
D A is false, but R is true.
- Q77.** A student takes Na_2CO_3 powder in a test tube and pours some drops of acetic acid over it. He observes:. **1 Mark**
- A** No reaction in the test tube. **B** Colourless gas with pungent smell.
C Bubbles of a colourless and odourless gas. **D** White fumes with smell of vinegar.
- Q78.** Calcium phosphate is present in tooth enamel. Its nature is: **1 Mark**
- A** Basic. **B** Acidic. **C** Neutral. **D** Amphoteric.
- Q79.** Which of the following is used for dissolution of gold? **1 Mark**
- A** Hydrochloric acid **B** Sulphuric acid **C** Nitric acid **D** Aqua regia
- Q80.** A student requires hard water for an experiment in his laboratory which is not available in the neighbouring area. In the laboratory, there are some salts, which when dissolved in distilled water can convert it into hard water. Select from the following groups of salts, a group, each salt of which when dissolved in distilled water will make it hard. **1 Mark**
- A** Sodium chloride, Potassium chloride. **B** Sodium sulphate, Potassium sulphate.
C Sodium sulphate, Calcium sulphate. **D** Calcium sulphate, Calcium chloride.
- Q81.** The indicator which produces a pink colour in an alkaline solution is: **1 Mark**
- A** Methyl orange. **B** Turmeric paper. **C** Phenolphthalein. **D** Litmus paper.
- Q82.** The property which is common between vinegar and curd is that they: **1 Mark**
- A** Have sweet taste. **B** Have bitter taste. **C** Are tasteless. **D** Have sour taste.
- Q83.** Select from the following the best experimental set-up for tracing the path of a ray of light through a glass slab: **1 Mark**
- 
- A** I **B** II **C** III **D** IV
- Q84.** The colour of raisins as used in the experiment, 'to determine the percentage of water absorbed by raisins' was. **1 Mark**
- A** White. **B** Yellow. **C** Dark brown. **D** Pink.
- Q85.** If the pH value of a solution is greater than 7, solution is: **1 Mark**
- A** Basic **B** Acidic **C** Neutral **D** Cannot be predicted
- Q86.** Hard water required for an experiment is not available in a school laboratory. However, following salts are available in the laboratory. Select the salts which may be dissolved in water to make it hard for the experiment. **1 Mark**
1. Calcium Sulphate.

2. Sodium Sulphate.
3. Calcium Chloride.
4. Potassium Sulphate.
5. Sodium Hydrogen Carbonate.
6. Magnesium Chloride.

A 1, 2 and 4 **B** 1, 3 and 6 **C** 3, 5 and 6 **D** 2, 4 and 5

Q87. Which of the following gives the correct increasing order of acidic strength? **1 Mark**

- A** Water < acetic acid < hydrochloric acid. **B** Water < hydrochloric acid < acetic acid.
C Acetic acid < water < hydrochloric acid. **D** Hydrochloric acid < water < hydrochloric acid.

Q88. A student adds 4 ml of acetic acid to a test tube containing 4 mL of distilled water. He then shakes the test tube and leaves it to settle. After about 10 minutes he observes: **1 Mark**

- A** A layer of water over the layer of acetic acid. **B** A layer of acetic acid over the layer of water.
C A precipitate settling at the bottom of the test tube. **D** A clear colorless solution.

Q89. Which among the following is not a base? **1 Mark**

- A** NaOH. **B** KOH. **C** NH₄OH. **D** C₂H₅OH.

Q90. You are asked to prepare hard water in your laboratory. Select a group from the following groups of salts, any salt of which you may dissolve in distilled water to obtain hard water. **1 Mark**

- A** NaCl; Na₂SO₄; KCl **B** NaCl; CaCl₂; KCl **C** CaCl₂; CaSO₄; MgSO₄ **D** Na₂SO₄; CaSO₄; MgSO₄

Q91. Sodium hydrogen carbonate when added to acetic acid evolves a gas. Which of the following statements are true about the gas evolved? **1 Mark**

1. It turns lime water milky.
2. It extinguishes a burning splinter.
3. It dissolves in a solution of sodium hydro.
4. It has a pungent odour.

- A** (i) and (ii) **B** (i), (ii) and (iii) **C** (ii), (iii) and (iv) **D** (i) and (iv)

Q92. At what temperature is gypsum heated to form Plaster of Paris? **1 Mark**

- A** 90°C **B** 100°C **C** 110°C **D** 120°C

Q93. For question two statements are given-one labelled Assertion (A) and the other labelled Reason (R). Select the correct answer to these questions from the codes (a), (b), (c) and (d) as given below: **1 Mark**

Assertion: pH = 7 signifies pure water.

Reason: pH of acetic acid is greater than 7.

- A** Both A and R are true, and R is correct explanation of the assertion.
B Both A and R are true, but R is not the correct explanation of the assertion. **C** A is true, but R is false.
D A is false, but R is true.

Q94. Corrosive effect of skin is caused by: **1 Mark**

- A** Acids **B** Bases **C** Salts **D** Water

Q95. What is formed when zinc reacts with sodium hydroxide? **1 Mark**

- A** Zinc hydroxide and sodium. **B** Sodium zincate and hydrogen gas.
C Sodium zinc-oxide and hydrogen gas. **D** Sodium zincate and water.

Q96. Chemical formula of baking soda is: **1 Mark**

- A** MgSO₄ **B** Na₂CO₃ **C** NaHCO₃ **D** MgCO₃

Q97. For question two statements are given-one labelled Assertion (A) and the other labelled Reason (R). Select the correct answer to these questions from the codes (a), (b), (c) and (d) as given below: **1 Mark**

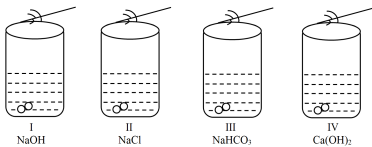
Assertion: Acetic acid does not act as an acid in benzene solution.

Reason: Benzene is non-polar.

- A** Both A and R are true, and R is correct explanation of the assertion.

- B** Both A and R are true, but R is not the correct explanation of the assertion.
- C** A is true, but R is false.
- D** A is false, but R is true.

Q98. A student added acetic to test tubes I, II, III and IV containing the labelled substance and then brought a burning splinter near the mouth of each test tube. **1 Mark**



The splinter would be extinguished when brought near the mouth of test tube:

- A** I
- B** II
- C** III
- D** IV

Q99. The acid present in the body of red ant is _____. **1 Mark**

- A** Formic acid
- B** Acetic acid
- C** Nitric acid
- D** Sulphuric acid

Q100. Acetic acid was added to a solid X kept in a test tube. A colourless and odourless gas was evolved. The gas was passed through lime water which turned milky. It was concluded that. **1 Mark**

- A** Solid X is sodium hydroxide and the gas evolved is CO₂.
- B** Solid X is sodium bicarbonate and the gas evolved is CO₂
- C** Solid X is sodium acetate and the gas evolved is CO₂
- D** Solid X is sodium chloride and the gas evolved is CO₂